

Questions with Answer Keys

MathonGo

Q1: 16 March (Shift 2) - Numerical

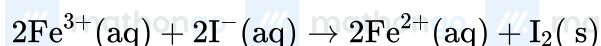
A 5.0 mol dm^{-3} aqueous solution of KCl has a conductance of 0.55 mS when measured in a cell constant 1.3 cm^{-1} . The molar conductivity of this solution is _____ $\text{mS m}^2 \text{ mol}^{-1}$. (Round off to the Nearest Integer)

Q2: 17 March (Shift 2) - Numerical

A KC solution of conductivity 0.14 S m^{-1} shows a resistance of 4.19Ω in a conductivity cell. If the same cell is filled with an HCl solution, the resistance drops to 1.03Ω . The conductivity of the HCl solution is _____ $\times 10^{-2} \text{ S m}^{-1}$. (Round off to the Nearest Integer).

Q3: 18 March (Shift 1) - Numerical

For the reaction



the magnitude of the standard molar free energy change, $\Delta_r G_m^\circ = -$ _____ kJ (Round off to the Nearest Integer).

Q4: 18 March (Shift 2) - Numerical

The molar conductivities at infinite dilution of barium chloride, sulphuric acid and hydrochloric acid are $280, 860$ and $426 \text{ S cm}^2 \text{ mol}^{-1}$ respectively. The molar conductivity at infinite dilution of barium sulphate is _____ $\text{S cm}^2 \text{ mol}^{-1}$ (Round off to the Nearest Integer).

Questions with Answer Keys

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Answer Key

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Q1 (14.3)

Q2 (57)

Q3 (45)

Q4 (288)

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