

Questions with Answer Keys

MathonGo

Q1: 16 March (Shift 2) - Single Correct

Identify the elements X and Y using the ionisation energy values given below:

Ionization energy (kJ/mol)

	1 st	2 nd
X	495	4563
Y	731	1450

(1) $X = \text{Na}$; $Y = \text{Mg}$ (2) $X = \text{Mg}$; $Y = \text{F}$ (3) $X = \text{Mg}$; $Y = \text{Na}$ (4) $X = \text{F}$; $Y = \text{Mg}$

Q2: 16 March (Shift 2) - Single Correct

The characteristics of elements X , Y and Z with atomic numbers, respectively, 33, 53 and 83 are:(1) X and Y are metalloids and Z is a metal.(2) X is a metalloid, Y is a non-metal and Z is a metal.(3) X , Y and Z are metals.(4) X and Z are non-metals and Y is a metalloid

Q3: 17 March (Shift 1) - Single Correct

The absolute value of the electron gain enthalpy of halogens satisfies:

(1) $\text{I} > \text{Br} > \text{Cl} > \text{F}$ (2) $\text{Cl} > \text{Br} > \text{F} > \text{I}$ (3) $\text{Cl} > \text{F} > \text{Br} > \text{I}$ (4) $\text{F} > \text{Cl} > \text{Br} > \text{I}$

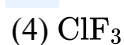
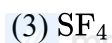
Q4: 17 March (Shift 1) - Single Correct

Which of the following compound CANNOT act as a Lewis base?

(1) NF_3

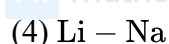
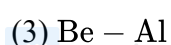
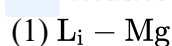
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Q5: 17 March (Shift 2) - Single Correct

The set of elements that differ in mutual relationship from those of the other sets is :



Q6: 18 March (Shift 1) - Single Correct

The ionic radius of Na^+ ions is 1.02 \AA . The ionic radii (in \AA) of Mg^{2+} and Al^{3+} , respectively, are

(1) 1.05 and 0.99

(2) 0.72 and 0.54

(3) 0.85 and 0.99

(4) 0.68 and 0.72

Q7: 18 March (Shift 2) - Single Correct

The first ionization energy of magnesium is smaller as compared to that of elements X and Y, but higher than that of Z. The elements X, Y and Z, respectively, are :

(1) chlorine, lithium and sodium

(2) argon, lithium and sodium

(3) argon, chlorine and sodium

(4) neon, sodium and chlorine

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Answer Key

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Q1 (1)

Q2 (2)

Q3 (3)

Q4 (2)

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Q5 (4)

Q6 (2)

Q7 (3)

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