

Master Math for JEE Main & JEE Advanced

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For JEE Main 2020 April



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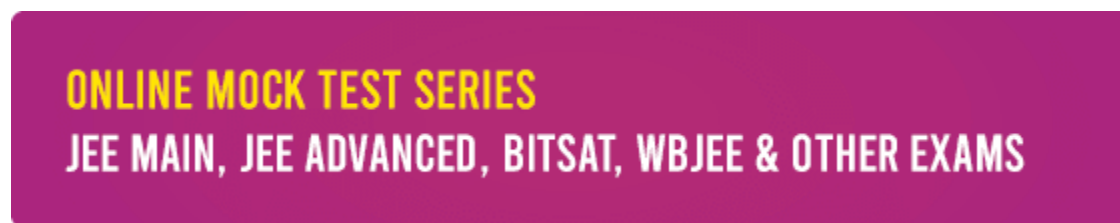
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JEE Mains 2020 Jan Chapter wise Question Bank

Surface Chemistry

Q1

Flocculation value for As_2S_3 sol by HCl is 30 m mole L^{-1} . Calculate mass of H_2SO_4 required in gram for 250 mL sol.

7th Jan Evening

Sol

00.37

For 1L sol 30 m mol of HCl is required

\therefore For 1L sol 15 m mol H_2SO_4 is required

For 250 mL of sol

$$\frac{15}{4} \times 10^{-3} \text{ m mol } \text{H}_2\text{SO}_4 \equiv 0.3675 \text{ g}$$

Q2

According to Hardy Schultz rule, correct order of flocculation value for $\text{Fe}(\text{OH})_3$ sol is :

$\text{Fe}(\text{OH})_3$ sol के लिए हार्डी-शुल्ज नियम के अनुसार ऊर्णन मान का सही क्रम है—

- (1) $\text{K}_2\text{CrO}_4 > \text{K}_3[\text{Fe}(\text{CN})_6] > \text{KNO}_3 > \text{KBr} = \text{AlCl}_3$
- (2) $\text{K}_3[\text{Fe}(\text{CN})_6] > \text{K}_2\text{CrO}_4 > \text{KNO}_3 = \text{KBr} = \text{AlCl}_3$
- (3) $\text{K}_3[\text{Fe}(\text{CN})_6] < \text{K}_2\text{CrO}_4 < \text{KNO}_3 = \text{KBr} = \text{AlCl}_3$
- (4) $\text{KNO}_3 > \text{KBr} = \text{K}_2\text{CrO}_4 > \text{AlCl}_3 = \text{K}_3[\text{Fe}(\text{CN})_6]$

8th Jan Morning

Sol

(3)

According to hardy-schultz rule,

$$\text{Coagulation value or flocculation value} \propto \frac{1}{\text{Coagulation power}}$$

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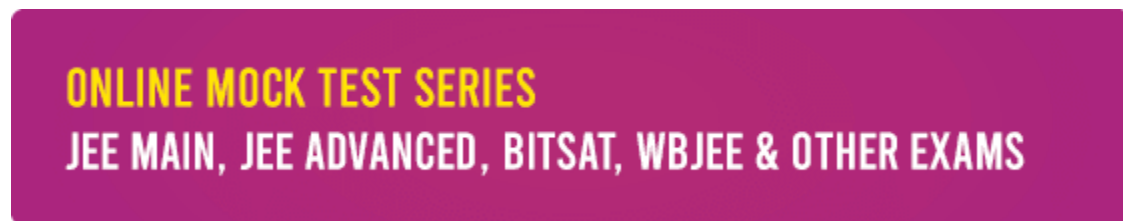
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