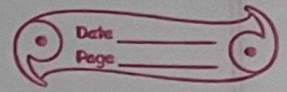


Periodic Table

Part - II



Periodic Properties (cont.)

4. Chemical reactivity

across a period: first decreases then increases
down the group: chemical reactivity increases

* Caesium is the most reactive metal

* Fluorine is the most reactive non-metal

5. Ionisation Potential

Ionisation potential depends on:

i) Atomic size: greater is the atomic size,
lesser the force of attraction,
 \therefore less ionisation energy.

ii) Nuclear charge: greater is the nuclear charge,
greater is the ionisation energy

trends in ionisation potential:
across the period: increases
down the group: decreases.

6. Electron Affinity:

Electron affinity depends on:

- i) Atomic size: Smaller the atomic size greater the electron affinity
- ii) Nuclear charge: Greater the nuclear charge greater is the electron affinity.

Trends in electron affinity:
Across a period: increases
Down the group: decrease

7. Electronegativity

Electronegativity depends on:

- i) Atomic size: Greater the atomic size less the electronegativity
- ii) Nuclear charge: Greater the nuclear charge greater the electronegativity

Trends in electronegativity:
Across a period: increases
Down the group: decreases