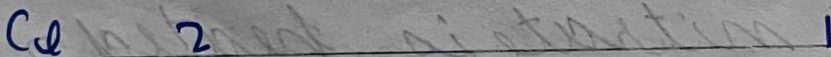
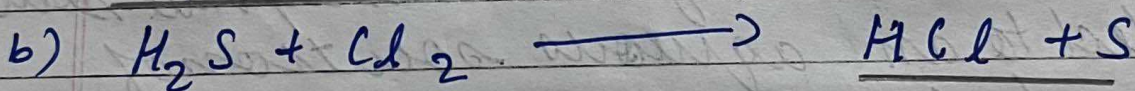


6. Complete and balance the following chemical equations:



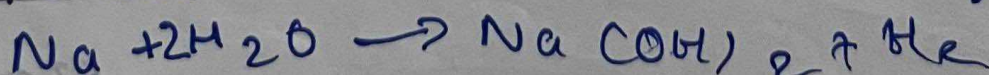
∴ The balanced chemical equation is $N_2 + O_2 \longrightarrow 2NO$



∴ The balanced chemical equation is



∴ The balanced chemical equation is

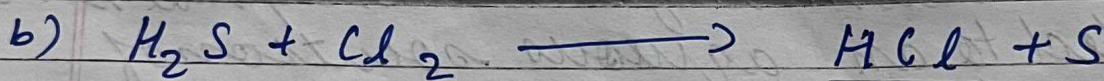


ans. When dilute sulphuric acid is added to barium chloride solution, white precipitate is formed.

6. Complete and balance the following chemical equations:



∴ The balanced chemical equation is $N_2 + O_2 \longrightarrow 2NO$



∴ The balanced chemical equation is



∴ The balanced chemical equation is $Na + 2H_2O \longrightarrow Na(OH)_2 + H_2$



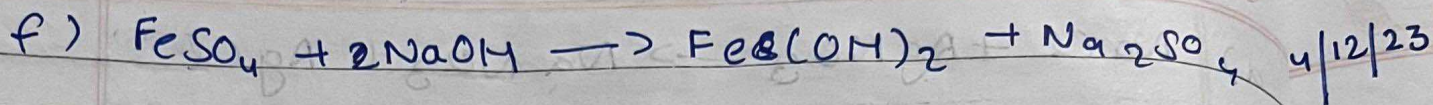
Na	1	1	
Cl	1	1	
Ag	1	1	
N	1	1	
O	3	3	

\therefore The balanced chemical equation is $\text{NaCl} + \text{AgNO}_3 \longrightarrow \text{NaNO}_3 + \text{AgCl}$



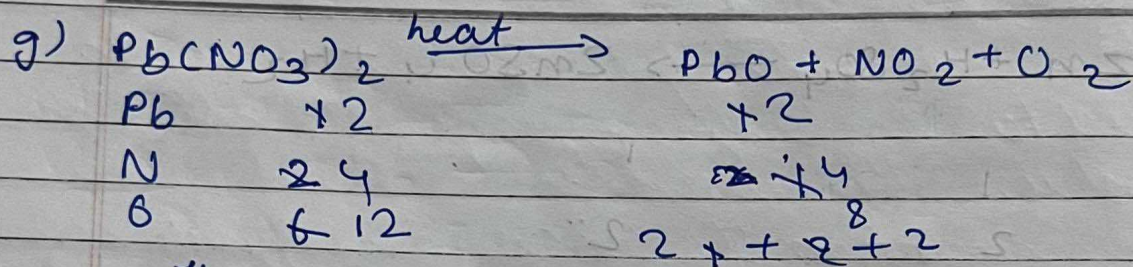
Zn	1	1
H	2	2
S	1	1
O	4	4

\therefore The balanced chemical equation is $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$

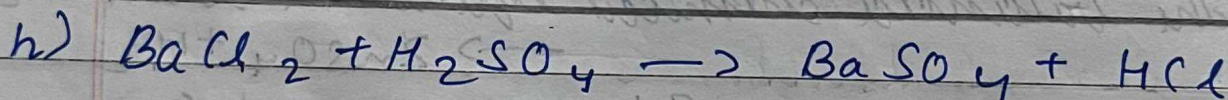


Fe	1		1
S	1		1
O	4 + 1		2 + 4
Na	1		2
OH	1		2

∴ The balanced chemical equation is
 $\text{FeSO}_4 + 2\text{NaOH} \rightarrow \text{Fe(OH)}_2 + \text{Na}_2\text{SO}_4$



∴ The balanced chemical equation is
 $2\text{Pb(NO}_3)_2 \rightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$

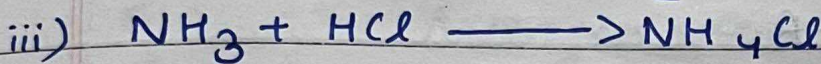


Ba	1		1
Cl	2		+ 2
H	2		+ 2
S	1		1
O	4		4

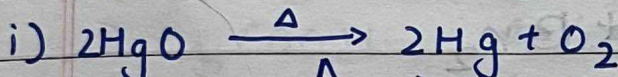
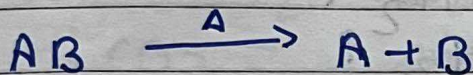
∴ The balanced chemical equation is
 $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{HCl}$

Types of Chemical Reaction

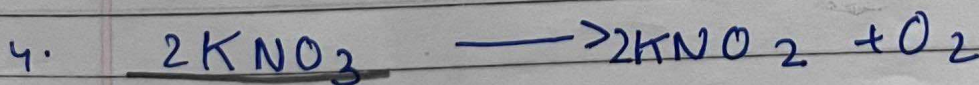
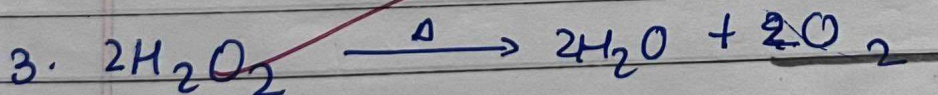
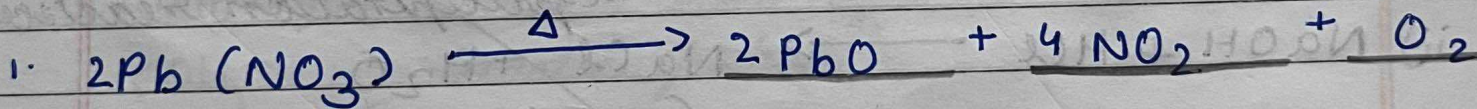
1. Combination reaction

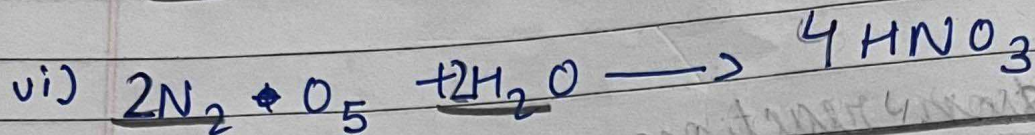


2. Decomposition reaction

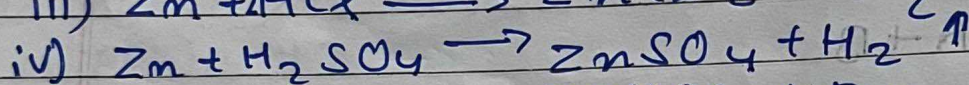
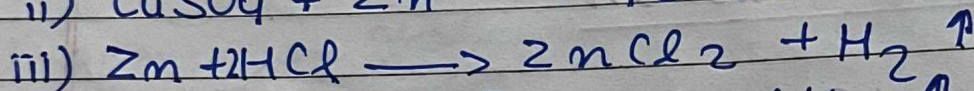
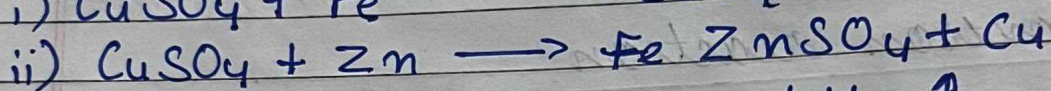
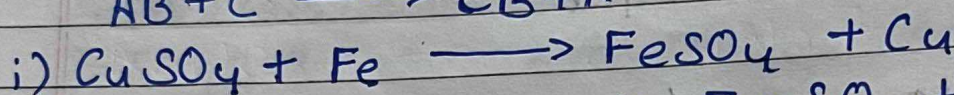


8. Complete and balance the following chemical reaction.

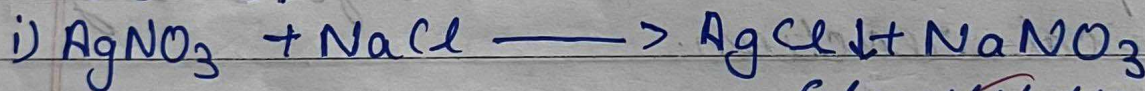




3. Displacement reactions



4. Double displacement reaction



(precipitation reaction)



(neutralization)