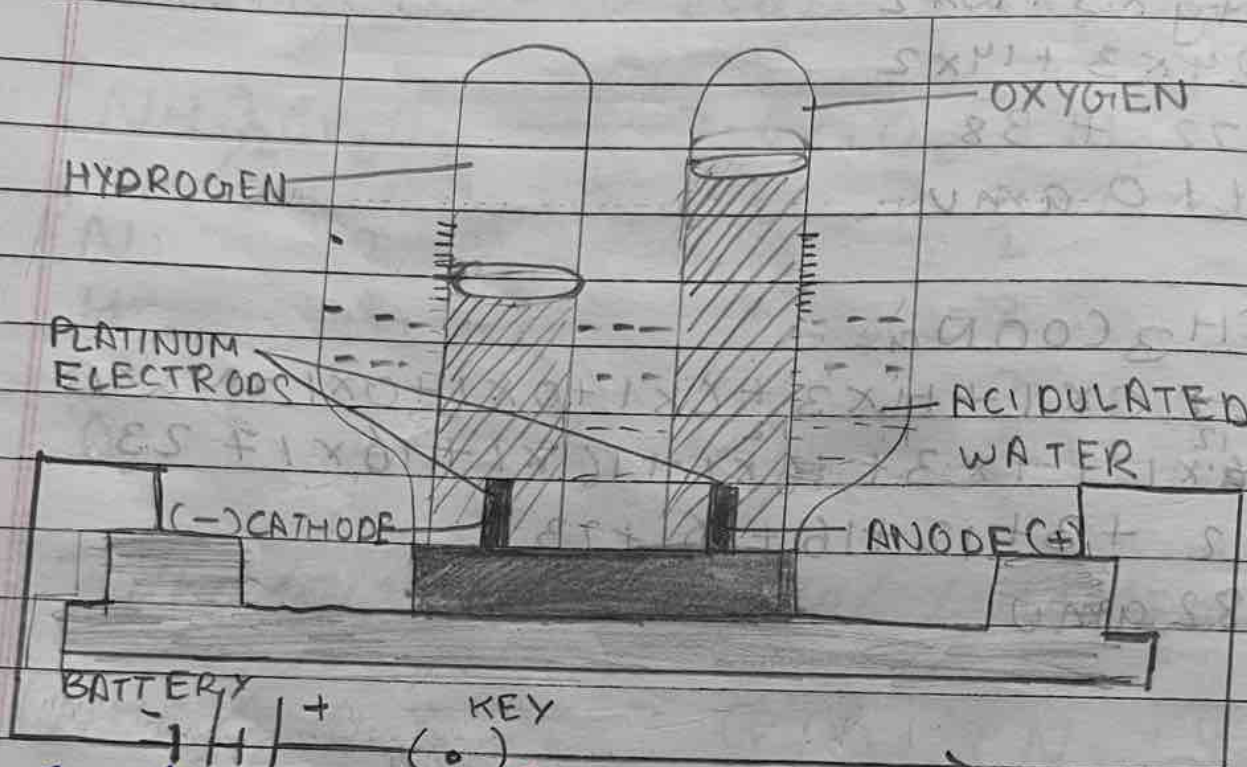
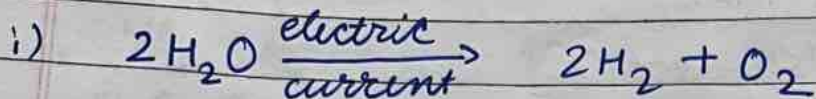


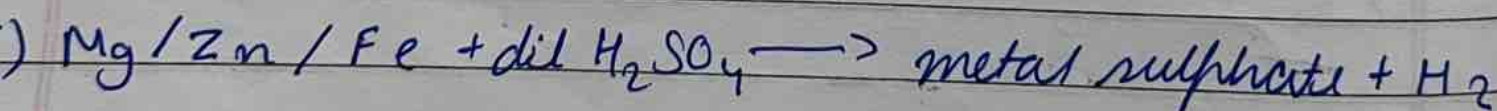
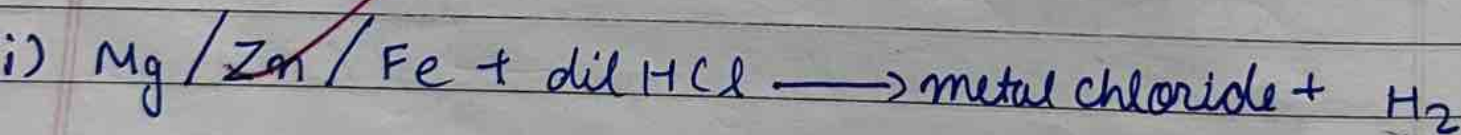
Hydrogen

Preparation of Hydrogen

i) By Electrolysis of water

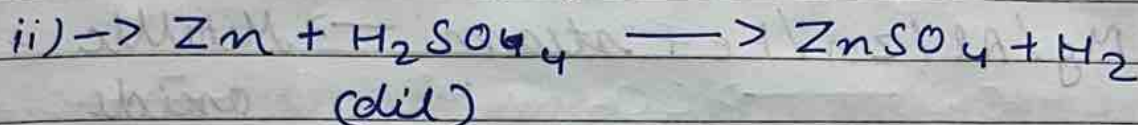
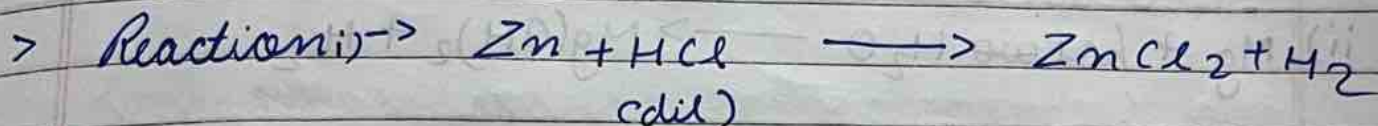


ii) By the action of dilute acid on active metals

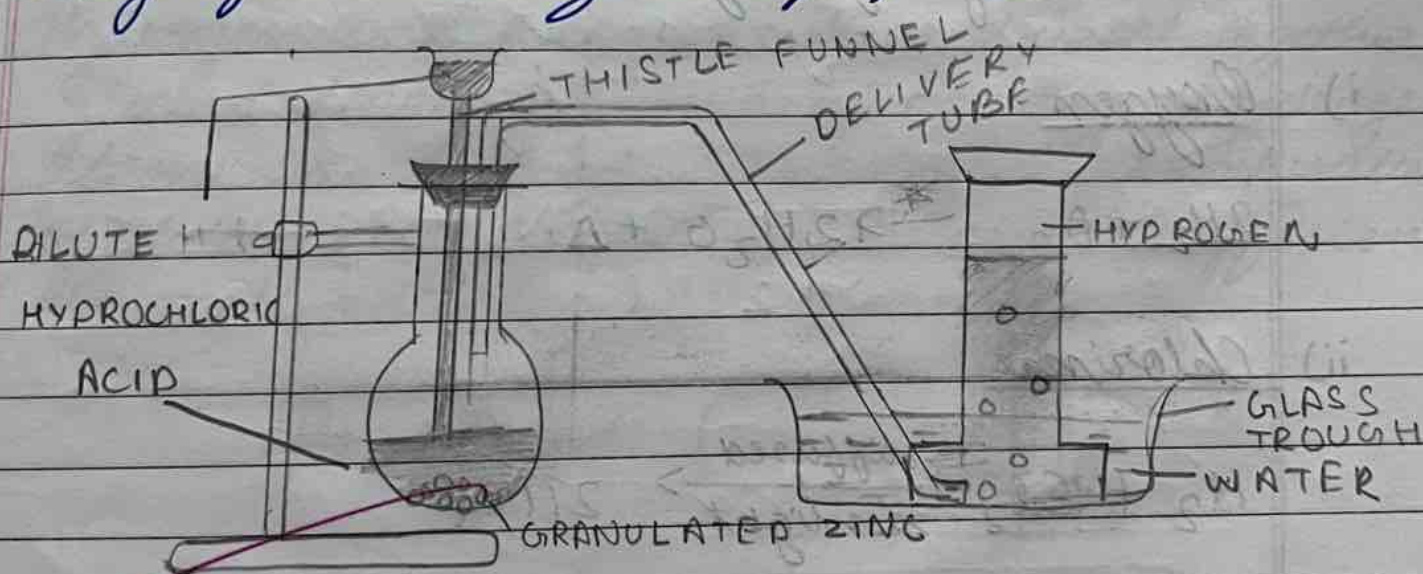


iii) By laboratory method preparation of Hydrogen

> Reactants: dil HCl / H₂SO₄, granulated zinc

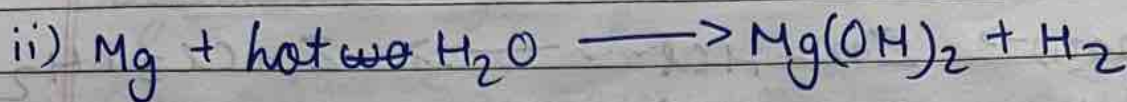
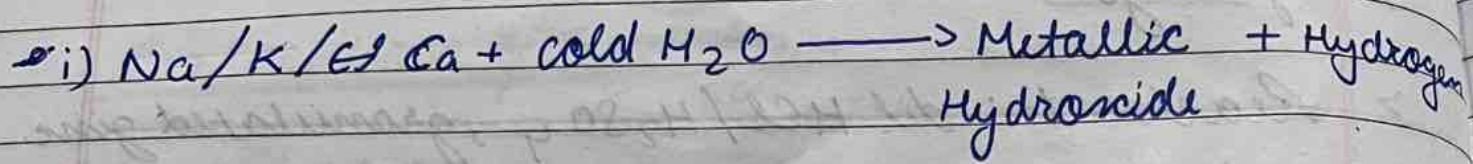


⊗ Why is granulated zinc is preferred



→ method of collection - downward displacement of water

Action of water or steam on metals

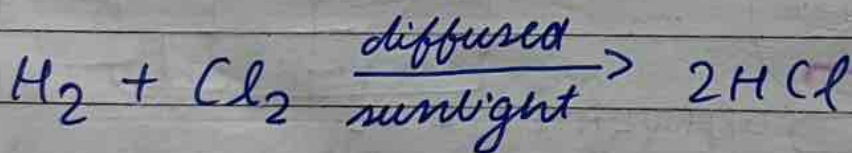


Reaction of Hydrogen with non metals

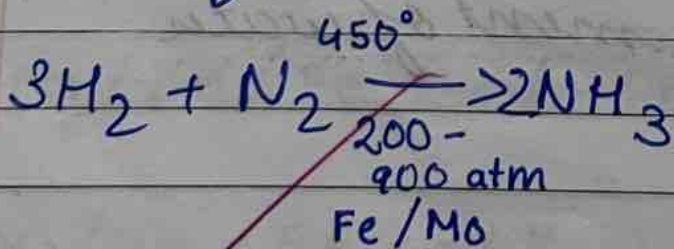
i) Oxygen



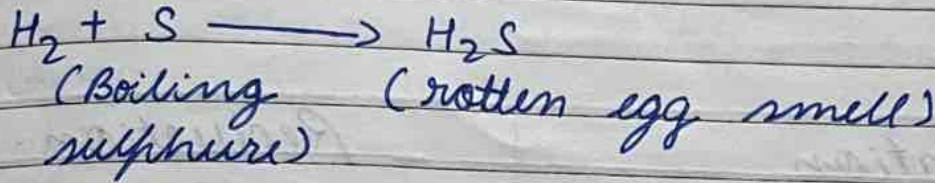
ii) Chlorine



iii) Nitrogen



iv) Sulphur



Q Give reason for the following :-

1. Hydrogen can be used for cutting and welding metals

Hydrogen and oxygen when burnt together gives a flame, known as oxy-hydrogen flame and it has a high temperature of about 2800°C - 3000°C , which is sufficient to melt metals.

~~For Oxyd~~